

## Le Chatlier's Principle

Name: \_\_\_\_\_

Date: \_\_\_\_\_ Per: \_\_\_\_\_

### Part 1: Le Chatlier's Principle Demonstration

**Directions:** Fill in the table with observations and explanations as you watch the demonstration of Le Chatlier's Principle.

Demonstration Description	Reaction	Observation	Explanation
Demonstration 1: Effect of the addition of a reactant on the equilibrium position	$\text{Co}(\text{H}_2\text{O})_6^{2+} + 4\text{Cl}^{-} + \text{HEAT} \rightleftharpoons \text{CoCl}_4^{2-} + 6\text{H}_2\text{O}$ (pink) (blue)		
Demonstration 2: Effects of addition of a product on equilibrium position	$\text{Co}(\text{H}_2\text{O})_6^{2+} + 4\text{Cl}^{-} + \text{HEAT} \rightleftharpoons \text{CoCl}_4^{2-} + 6\text{H}_2\text{O}$ (pink) (blue)		
Demonstration 3: Effect of temperature on equilibrium position	$\text{Co}(\text{H}_2\text{O})_6^{2+} + 4\text{Cl}^{-} + \text{HEAT} \rightleftharpoons \text{CoCl}_4^{2-} + 6\text{H}_2\text{O}$ (pink) (blue)		
Demonstration 4: Removal of a reactant on the equilibrium position	$\text{Co}(\text{H}_2\text{O})_6^{2+} + 4\text{Cl}^{-} + \text{HEAT} \rightleftharpoons \text{CoCl}_4^{2-} + 6\text{H}_2\text{O}$ (pink) (blue)		